IUPAC Task Group on Atmospheric Chemical Kinetic Data Evaluation

 – Data Sheet AQ\_OH\_43

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**HO(aq) + HO(CH2)4OH(aq) → products**

**Rate coefficient data**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| k/ L mol-1 s-1 | T/K | pH | I/ mol L-1 | Reference | Technique/ Comments |
| *Relative Rate Coefficients* |
| 3.23  109 | 294 | - | - | Adams et al., 1965 | PR / UV-Vis(a) |
| (3.4 ± 0.1) × 109 | 298 | 7 | - | Hoffmann et al., 2009 | LFP-LPA (b) |
| 9.7 × 1010 exp[-(1000 ± 80)/T] | 288 - 328 | 7 | - |  | LFP-LPA (b1) |

*GR* (aq): Aqueous phase thermochemical data not available. As well, gas phase thermochemical data *R* (g) are not available.

**Comments**

(a) Reference reaction: HO + SCN- with *k*(HO + SCN-) = 6.6  109 M-1s-1; the selected reference rate coefficient *k* = 1.10 × 1010 M‑1s‑1 was used for recalculation; No exact value is given for the initial concentrations of the reactants; pH is given as natural; as no exact temperature is given, T = 294 K is assumed for room temperature.

(b) Radicals generated by laser flash photolysis (LFP) of H2O2 (*c*(H2O2) = 1  10-4 M) at 248 nm; reference reaction: HO + SCN- with *k*(HO + SCN-) = 1.24 × 1010 M‑1 s‑1; the selected reference rate constant *k* = 1.19 × 1010 M‑1s‑1 was used for recalculation; *c*(KSCN) = 1.59  10‑5 M. Arrhenius expression (b1) was calculated using the recalculated experimental data from Hoffmann et al. (2009), based on the recommended Arrhenius expression *k*(T) = 3.45209 × 1012 × exp(-14050 / RT) (Zhu et al., 2003).

**Preferred Values**

|  |  |  |
| --- | --- | --- |
| **Parameter** | **Value** | ***T*/K** |
|  |  |  |
| *k* / L mol-1 s-1 | 3.33 × 109  | 298 |
|  |  |  |
| *k* / L mol-1 s-1 | 9.04 × 1010 exp[-(1000)/T] | 288 – 328 |
|  |  |  |

*Reliability*

|  |  |  |
| --- | --- | --- |
| Δ log *k* | ±0.09 | 298 |
| Δ EA/R | ±70 | 288 – 328 |
|  |  |  |

*Comments on Preferred Values*

For the recommended rate constant, data determined by Adams et al. (1965) and Hoffmann et al. (2009) have been used for regression. The available data agree well with each other as well as with the former recommendation by Buxton et al. of 1988 when the latter is recalculated to 3.2 × 1010 M‑1 s‑1. The uncertainty is estimated as ±20% or Δ log *k* = ±0.09.

**References**

Adams, G. E., Boag, J. W. and Michael, B. D.: Trans. Faraday Soc., 61, 1417-1424, 1965.

Buxton, G. V., Greenstock, C. L., Helman, W. P. and Ross, A. B.: J. Phys. Chem. Ref. Data, 12(2), 513 – 886, 1988.

Chin, M., and Wine, P. H: J. Photochem. Photobiol., A, 69(1), 17-25, 1992.

Hoffmann, D., Weigert, B., Barzaghi, P. and Herrmann, H.: Phys. Chem. Chem. Phys., 11, 9351-9363, 2009.

Monod, A. and Doussin, J. F.: Atmos. Environ., 42, 7611–7622, 2008.

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T-dependent rate constants for the reaction of butane-1,4-diol with HO in aqueous solution. All data shown in the figure has been used for regression.